Lighter Than Air HW

Read and outline (Cornell style) **Section 14.4** in your chemistry textbook. ****Section outlines must be at least ½ page with 3-4 topic questions.**** Then answer the following assessment questions.

1. How many grams of CO_2 are in a 1.0-L balloon at STP?
2. What volume (in mL) will 0.00922g of H_2 gas occupy at STP?
3. Consider the following reaction: $C_3H_8 + O_2 \rightarrow CO_2 + H_2O$
    a. Balance the equation.
    b. How many liters of propane gas will undergo complete combustion with 34.0 L of oxygen gas?
4. Consider the following reaction: $H_2 + O_2 \rightarrow H_2O$
    a. Balance the equation.
    b. Determine the volume of hydrogen gas needed to react completely with 5.00 L of oxygen gas to form water.
5. Solid potassium metal will react with Cl_2 gas to form potassium chloride. How many liters of Cl_2 gas are needed to completely react with 0.204g of potassium at STP?